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| Student check list: 🗹 when you know…  The learning goals and success criteria for this term  Changes to routines e.g. excursions | | | | Assessment due dates | |
| When assessment practice lessons will occur (exemplars)  When revision lessons will occur | |
| **WK** | **Wk. Beg** | **Holidays or variations this week** | **Lesson 1** | **Lesson 2** | **Lesson 3** |
| 1 | 8 Jul. 24 | **Work Exp.** |  |  |  |
| 2 | 15 Jul. 24 |  | **Assumed knowledge:**  Atomic Structure, Chemical Formulas, Balanced equations, General structure of the Periodic Table | **Percent composition** of elements in a chemical formula | **Mole concept:**  Avogadro’s number  Molar Mass  Formula to determine moles |
| 3 | 22 Jul. 24 |  | **Mole concept:**  Solve basic problems involving mass, moles and molar mass | **Mole concept:**  Use the mole concept to calculate mass of reactants/products in a balanced equation | **Mole concept:**  Use the mole concept to calculate mass of reactants/products in a balanced equation |
| 4 | 29 Jul. 24 |  | **Mole concept:**  Identifying limiting reagents in a balanced chemical equation | **Mole concept:**  Solve problems using limiting reagents to determine the mass of products formed | **Aqueous solutions:**  Understand the term concentration (Molarity) and explain how different molarity solutions are made. |
| 5 | 5 Aug. 24 |  | **Aqueous solutions:**  Solve problems involving concentrations and volumes of reactants to determine amount of products | **Aqueous solutions:**  Solve problems involving concentrations and volumes of reactants to determine amount of products | **Measurement uncertainty and error:**  Understand the role of significant figures in calculations |
| 6 | 12 Aug. 24 |  | **Measurement uncertainty and error:**  Qualitative and Quantitative data  Random and systemic errors  Measurement uncertainties | **Measurement uncertainty and error:**  Calculate the measurement uncertainties in processed data using absolute uncertainties and percentage uncertainties | **Student Experiment:**  Overview of what a student experiment is.  What a report looks like.  Scaffolding that will be provided and timelines for work to be completed |
| 7 | 19 Aug. 24 |  | **Class practical:**  Reaction of hydrochloric acid and Sodium Thiosulfate (Hypo) | **Student Experiment:**  Research question  Rationale for the experiment  Original experiment  Modifications to the methodology | **Student Experiment:**  Management of risks  How the raw data will be collected |
| 8 | 26 Aug. 24 | **Friday (30th)**  **Pupil free day** | **Student Experiment:**  Perform experiment | **Student Experiment:**  Perform experiment | **STUDENT FREE DAY** |
| 9 | 2 Sep. 24 |  | **Student Experiment:**  Data processing | **Student Experiment:**  Trends, patterns and relationships | **Student Experiment:**  Evaluation of the methodology  Suggestions for improvements |
| 10 | 9 Sep. 24 |  | **Student Experiment:**  Conclusion  Reference list |  |  |